

Name \_\_\_\_\_

Ch 5 and 6 Test Study Guide CP Living

Define the following terms:

Polyatomic ion

Non-electrolyte

Covalent Bond

Soluble

Ionic Bond

Insoluble

Crystal Lattice

Isomer

Electrolyte

Octet Rule

Explain the model of metallic bonding and how metallic bonds results in the properties metal have.

Explain how ions bond to form a crystal lattice. What properties does the lattice give to ionic compounds?

Classify the following substances as having ionic, metallic, or covalent bonds:

LiCl \_\_\_\_\_

A nonelectrolyte \_\_\_\_\_

SiO<sub>2</sub> \_\_\_\_\_

CHCl<sub>3</sub> \_\_\_\_\_

Platinum Metal \_\_\_\_\_

Ca(NO<sub>3</sub>)<sub>2</sub> \_\_\_\_\_

Explain the kind of bonding taking place in Na<sub>2</sub>SO<sub>4</sub>

Complete the following table:

Element	# of Valence e <sup>-</sup>	# of e <sup>-</sup> needed for an octet	# of bonds formed	# of unshared electron pairs
N				
O				
C				
F				
H				

Draw Lewis Structures for the following molecules/ions



Draw the isomers of the following molecules:

